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atoms of different

elements have

different masses.

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(all oxygen atoms have the same mass; and the mass of oxygen is always different from every other element's mass) □

Compounds

contain atoms of more than one element. (ex: NaCl)

□ In a particular compound, atoms of different

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elements always.
combine in the
same way.

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Chapter 4: Atomic
Structure Notes

Flashcards | Quizlet

Name: _____ Period:

_____ Unit 2 Atomic

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Notes The Periodic

Table _____ is

credited for

creating the first

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version of the modern periodic table _____ is a chart showing all the elements arranged in columns in such a way that all the elements in a given column exhibit similar chemical properties The periodic table is organized in order

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periodic is

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Unit 2 - Atomic

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Name ...

Atomic Theory

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What does the

atomic theory

state? It States that

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all matter is made of extremely small particles called atoms. 2. Which country is most often credited with the discovery of modern science? Ancient Greece 3. Which philosopher originally thought up the atomic theory? Democritus 4.

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Unit 3 Atomic

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Notes The Periodic

Table ... - Add you

answers from

number 1 together

- Round your

answer to 2

decimal places □

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Structure
Practice: 1. A sample of Cesium is 75% ^{133}Cs , 20% ^{132}Cs , and 5% ^{134}Cs . What is the average atomic mass? 2.

Carbon-12 is 98% abundant and Carbon-14 is 2% abundant.

Name: Period: Unit
3 Atomic Structure-

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- Write the atomic mass at the bottom of the square. Step 2: Determine the number of protons, neutrons, and electrons in each element. Step 3: Create an atomic structure diagram for each element. Step 4: Draw the Bonding Structure

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Structure for each element. 5

B Boron 10.81 P =

5 N = 6 E = 5

Atomic Structure:

Bonding Structure:

B 5 P+ 6 no

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STRUCTURE

DIAGRAMS Step 1:

Complete the

squares for each

element by adding

the atomic number,

name and atomic

mass. - Write the

atomic number at

the top of the

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Structure. - Write the element's name under the symbol. -

Write the atomic mass at the bottom of the square.

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of how the
numbers of
subatomic particles
makes an atom
unique from any
other type of atom
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Structure Problems Answers.

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